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# BOOKS

Chemical Process Principles-Part 1, Material and Energy Balances. O. A. Hougen, K. M. Watson, and R. A. Ragatz. Second Edition. John Wiley & Sons, Inc., New York (1954). 525 pages. \$8.50.

Professor Ragatz has collaborated with the original authors of the Chemical Process Principles series to revise Part 1, "Material and Energy Balances." While retaining the many virtues of the original edition, the authors have increased its utility by the addition of discussions on stagewise extraction, ternary equilibria, and time lag in stirred vessels, and in recognition of the new era R. G. Taecker has contributed a brief section on the thermochemistry of nuclear reactions. Further, two new chapters are found in this edition, one devoted to mathematical procedures and conversion of units, the other to absorption. New methods of estimating physical properties and the latest thermal data are included.

Although the law of corresponding states and the concept of reduced properties are briefly treated in a prelude to the section on methods of estimated critical properties, the generalized compressibility-factor chart is not discussed. Thus the authors proceed from a treatment of ideal-gas behavior directly to reduced and critical properties, including a method of estimating critical compressibility factor, without reference to the most common application of the law of corresponding states. Continuity would seem to be sacrificed here. For example, an apparently promising method of predicting latent heats of vaporization is later in the book presented in terms of reduced properties and gas-liquid compressibility factors, the significance and limitations of which are lost to the student in his ignorance of the generalized compressibility chart. The merits of these chapters for reference purposes are obvious; however, it is the reviewer's experience that the pedagogic value of these portions of the text would have been enhanced by the early introduction of the generalized Z chart. The semimythical character of the ideal-gas concept would thus have been impressed upon the student's mind, and the value and intrinsic limitations of the law of corresponding states could have been made apparent in the light of the authors' discussion of the Lennard-Jones model and molecular polarity. The student must learn as early as possible that chemical engineering tools are not de fide truths.

Viewed in its entirety, this edition is a good revision of a good textbook. The authors are to be commended.

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